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# The Role of Nanometer Scale Structure on Interfacial Energy

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## Abstract

Solid-liquid interfaces are of central importance in controlling processes that span fields from biology to quantum mechanics. At the nanoscale, surfaces are often structured with domains; this is especially true in biological systems, where folded proteins present surfaces with nanometer-scale domains of varying hydrophobicity and polarity. It has been established that the nanostructuring of such surfaces determines interfacial interactions and influences protein function, but an accepted framework providing a global and quantitative understanding is still lacking. The conventional continuum thermodynamic approach of capillary theory treats the interfacial energy  $\gamma_{SL}$  of surfaces that contain chemically distinct domains solely as a function of average surface composition, and hence it ignores structure. Here we show that, when a surface has domains commensurate in size with solvent molecules, its total  $\gamma_{SL}$  is determined not only by its average composition but also by a structural component that causes it to deviate from the continuum prediction by a substantial amount; in our system, by as much as 20%. Our studies of  $\gamma_{SL}$  are based on macroscopic contact angles as well as nanoscopic atomic force microscopy measurements that validate each other. By contrasting surfaces coated with mixed self-assembled monolayers that spontaneously organize into either molecular scale stripes (0.5 to 2 nm) or larger scale domains (>5 nm), we find that while the latter surfaces show the linear dependence of  $\gamma_{SL}$  on surface composition expected from continuum thermodynamics, surfaces with small domains show a non-monotonic trend strongly deviating from the linear behavior. Using molecular dynamics simulations, we show how the organization and degree of bonding of the solvent molecules at the interface is determined by the nano-structured surface, which in turn appreciably modifies  $\gamma_{SL}$ . We argue that these findings are general and demonstrate the important role of nanometer-scale structure on interfacial phenomena.

Interfacial phenomena are essential for processes such as protein dynamics<sup>1</sup>, heterogeneous catalysis<sup>2</sup>, biological surface activity<sup>3, 4</sup>, biosensing<sup>5</sup>, self-assembly<sup>6</sup>, and electron transfer<sup>7</sup>. One of the most striking examples of the importance of interfaces can be found in molecular biology. Protein folding and dynamics in aqueous solution rely on a well-controlled interface between the protein surface and its surrounding medium. It is postulated that the unique sequence of amino acids ensures optimized interactions with the solvent to support the protein's movement, activity and folding<sup>1, 8</sup>. However near the protein surface, the structure of the liquid is dynamic and intrinsically difficult to define, making experimental investigation challenging. Moreover, the surface structure of biological materials can itself depend on the local liquid environment<sup>8</sup>. Self assembled monolayers<sup>9, 10</sup> (SAMs) provide a unique opportunity to generate organic interfaces that can be used as models for the study of more complex, biological interfaces<sup>11</sup>. They can have compositions and surface structures that, to a first order, resemble biological materials' surfaces but (i) SAMs' structures do not reorganize<sup>12</sup> or denature<sup>13</sup> upon interaction with a solvent, and (ii) their composition can be easily modified, allowing for simple yet accurate systematic studies. SAMs on flat surfaces composed of mixtures of immiscible molecules spontaneously separate into domains of random shape and size (5 nm or larger, see Fig. 1 and Fig. 2A)<sup>12, 14</sup>. However, when SAMs of the same composition are formed on a gold nanoparticle<sup>15, 16</sup> in a given size range<sup>17, 18</sup> they separate into ordered, striped domains of alternating composition one or two molecules wide (i.e.  $\sim 5\text{\AA}$ , see Fig. 1)<sup>16, 19</sup>. In the nanoparticle case the domains are commensurate with (if not smaller than) the typical hydrophilic or hydrophobic domain size observed on folded proteins, while in the case of flat SAMs domains are considerably larger.

Interfacial energy  $\gamma_{SL}$  (S=solid, L=liquid, V=vapor) remains one of the most accessible measurable quantities used to characterize solid-liquid interfaces. It is directly linked to both medias' surface energies ( $\gamma_S$  and  $\gamma_L$  respectively) and to the work of adhesion ( $W_{SL}$ , i.e. the reversible free energy cost to separate the media), by the *Dupré* equation<sup>20</sup>:

$$\gamma_{SL} = \gamma_S + \gamma_L - W_{SL} \quad (1)$$

$W_{SL}$  and  $\gamma_{SL}$  describe the same phenomena. Experimentally, the contact angle  $\theta_{CA}$  formed by static, macroscopic liquid droplets at the surface of a flat solid in an inert atmosphere provides a direct measurement of  $W_{SL}$  through the *Young-Dupré* equation  $W_{SL} = \gamma_{LV} (1 + \cos \theta_{CA}) \approx \gamma_L (1 + \cos \theta_{CA})$ <sup>21</sup>. Current understanding of these interfaces is based on macroscopic observations and relies on the continuum thermodynamic approach (CTA). It is accepted that when a surface is made of different components its  $\gamma_{SL}$  is simply the weighted average of the  $\gamma_{SL}$  of the individual components<sup>22</sup> leading to a monotonic dependence of  $\theta_{CA}$  on composition. At a macroscopic scale, line tension, roughness<sup>23</sup>, and the presence of macroscopic gas pockets (e.g., air for the lotus leaf effect in Cassie's treatment<sup>24</sup>) all play a role in determining  $\gamma_{SL}$ <sup>25, 26</sup>. However, at the nanoscale, it is not known which surface parameters dominate  $\gamma_{SL}$  especially in the case of surfaces composed of chemically different domains. Multi-component surfaces in the CTA approach are either treated as uniform when domains of one component cannot be defined (e.g. single molecules dispersed in a matrix of other molecules) or as an ensemble of single component surfaces independently interacting with the liquid<sup>27</sup>. The range in domain size where one approximation ends and the other begins, or where either cannot be applied is unknown, yet it is accepted that solvents behave in unique ways when confined or in contact with nanoscale boundaries<sup>28, 29</sup>. Macroscopic  $\theta_{CA}$  measurements of mixed SAMs on flat surfaces ( $> 5$  nm domains) have shown that  $W_{SL}$ , and consequently  $\gamma_{SL}$ , exhibit a monotonic dependence on composition<sup>22, 30</sup>, therefore at this length scale the CTA remains valid. However, for mixed SAMs on nanoparticles ( $< 2$  nm domains) solubility experiments have reported a non-monotonic dependence of the nanoparticle saturation concentration on composition, hinting that this observation might be related to a failure of the representation of mixing enthalpy in CTA<sup>15</sup>; note that solubility is such a complex process depending on numerous factors that singling out any one of them is particularly challenging<sup>31</sup>. To the best of our knowledge, there has been to date no discussion of the role of surface structure in determining interfacial energy at solid-liquid interfaces. At the macroscale, the Cassie state of wetting (super-hydrophobicity) could appear to show such structure-dependent effects, but in reality trapped air pockets transform an apparent solid-liquid interface into a mixture of solid-air and solid-liquid interfaces. At the nanoscale, where true single interfaces

exist, the only available reports discussing structuring effects are solely theoretical and address solvent structure but not interfacial energy. Here, for the first time we present quantitative experimental evidence of an unanticipated and large dependence of  $W_{SL}$  (and hence  $\gamma_{SL}$ ) not only on composition but also on surface structure using nanopatterned SAMs with domain size commensurate with solvent molecules. We use molecular dynamics simulations and experimental evidences to corroborate a novel theory that proposes the existence of surface bound half cavities and structure modulated solvent density fluctuations at these interfaces.

To perform our studies we prepared surfaces coated with binary mixtures of 1-octanethiol (OT) and 6-mercaptohexan-1-ol (MHol) of varying ratios. To obtain  $\sim 5$  nm domains we assembled the molecules directly on flat gold surfaces by immersion in dilute ethanol solutions for one week at room temperature<sup>10</sup> (Fig. 1, Fig. 2A, SI). To generate  $<2$  nm domains we first synthesized mixed SAM-coated nanoparticles and then assembled smooth, multilayer films of nanoparticles using layer-by-layer assembly (Fig. 1, SI)<sup>32-34</sup>. Across compositions we found that the films had very similar thickness and every AFM image acquired had complete coverage of the substrate lacking pinhole defect (Fig. 2B-C, SI Table 1). ANOVA analysis of the roughness values obtained by AFM for slides of different compositions shows no statistically significant variation ( $\alpha=0.05$ , results summarized in table 1 of SI).

The  $W_{SL}$  for these films was measured using macroscopic  $\theta_{CA}$  and AFM techniques. Because the value for  $W_{SL}$  measured by  $\theta_{CA}$  effectively depends on the area close to the three-phase line<sup>35</sup>, this technique is intrinsically sensitive to film inhomogeneities or film reconstructions due to the pinning of the solvent droplet. To overcome this problem we relied on measurements performed in liquid with AFM, which allows direct and quantitative probing of local variations in the solvation layer near the surface down to the sub-nanometer level<sup>36-38</sup>, hence allowing for the measurement of interfacial energy in local regions away from the three phase line. Using Small Amplitude Modulation AFM, (SAM-AFM) operated in liquid at

amplitudes  $<2$  nm, we were able to simultaneously map the topography of the solid-liquid interface while measuring the energy dissipated by the tip into the interfacial liquid with molecular resolution (Fig. 2B-C, SI, Voitchovsky et al. submitted). Subsequent comparison of the measured local energy dissipation at the interface with that in the bulk liquid allowed us to estimate  $W_{SL}$  using a simple model (see SI). Although the AFM-based method bears significant uncertainty due to both experimental and modelling parameters, we always observed a qualitative, and statistically significant agreement with  $\theta_{CA}$  measurements of  $W_{SL}$  (Fig. 2D-E, and SI). Furthermore, whereas the uncertainty associated with AFM measurements is notable in absolute values of  $W_{SL}$ , also because, in the absence of calibration, this technique has a significant linear deviation relative to the CA measurements (see 1F and 2C in SI), such uncertainty is greatly reduced when studying variation of  $W_{SL}$  between different samples (nearly identical experimental conditions, see SI). Summarizing, the agreement between microscopic and nanoscopic measurements allows us to state unequivocally that the effects measured are due to specific nanoscale interactions between the SAMs and the liquid.

The most important results are summarized in Fig. 2D-E. The dependence of  $W_{SL}$  on average surface composition follows two markedly different trends when comparing flat and nanoparticle films. For flat SAMs (Fig. 2A, D), we find a monotonic dependence of  $W_{SL}$  on composition, as the CTA would predict. For the nanoparticle surfaces, where domains are not wider than 2 nm (Fig. 2B), the trend is clearly non-monotonic (Fig. 2E), and hence the CTA has to be considered inapplicable. In order to substantiate these observations we carried a comprehensive statistical analysis of our data summarized in section 2 of the SI, we find that both the AFM and CA trend for flat SAM are strictly monotonic ( $R^2$  0.413 and 0.940 respectively), while the trends for the nanoparticle surfaces are non-monotonic (ANOVA analysis on the residuals after quadratic fitting,  $\alpha < 0.01$  in both cases). The largest deviations from the expected linear behaviours are at the 1:2 and 2:1 compositions (20 and 12 mN/m respectively). In addition to the 1:1 composition (discussed below) these two are the compositions where the narrow (in this case  $<1.2$  nm wide) ribbon-like stripe morphology is

dominant<sup>16</sup>; hence these are the compositions where one might expect a stronger structure effect on interfacial energy. It is worth noting that while in both the flat SAM and the nanoparticle case the actual composition achieved may deviate slightly from the nominal composition used in the SAM preparation, and these deviations may be different for flat and nanoparticle surfaces, the trends are clearly different from each other and it is evident that both cannot be described by the same continuum model. The domain size and local ordering are the only notable differences between the two surfaces and yet the difference in  $W_{SL}$  is as large as 20%. The curvature of the nanoparticle cannot explain these observations since the root mean square (RMS) roughness is statistically the same for all mixed ligand NP films.

These data validate our assumption that a structured surface induces the organization of solvent (as probed by AM-AFM in liquid) whose molecular structure substantially affects  $W_{SL}$ , an intrinsically macroscopic property. We observe a non-monotonic dependence of  $W_{SL}$  on composition in water as well as in glycerol (see SI, Fig. S1), suggesting that this behaviour is general. Because we have performed the bulk of our experiments and simulations in water, for clarity of the discussion that follows we use terms specific to water (e.g. hydrophobic) without loss of generalization. We hypothesize that two molecular mechanisms, which we describe in detail below, compete to produce the observed non-monotonic behavior. To test our hypothesis, we carried out explicit, all-atom molecular dynamics (MD) simulations of striped NPs of different compositions (33%, 50% and 67% MHol, with stripe widths of 0.6, 0.7, and 1.2 nm, respectively), surrounded by water molecules (details in SI). The corresponding OT stripe widths are 1.2 nm, 0.7 nm and 0.6 nm, respectively. The stripe widths were chosen based on the average experimental stripe-widths for different compositions. For comparison, we also carried out simulations of ‘Janus’ NPs, on which the two components were completely separated rather than arranged into stripes (Fig. SI 7), but with identical compositions (Fig. S1 8). The simulations provide access to detailed molecular information to directly investigate both hypothesized mechanisms.



The first hypothesized molecular mechanism draws from the theory of cavitation energies for small, poorly soluble molecules<sup>39</sup>. Cavitation involves the encapsulation by solvent molecules of a hydrophobic solute within an intact but strained hydrogen-bonded cage,<sup>39-41</sup>. For a surface composed of hydrophobic and hydrophilic nanoscale domains, the ‘phobic’ molecules will push the solvent away from the surface while the ‘philic’ molecules will pull the solvent close to the surface. As illustrated in Fig. 3A, on macroscopic domains this should result in two markedly different average distances of the first layers of solvent molecules from the MHol and OT stripes, respectively; this is confirmed by density profiles obtained from our molecular dynamics simulations shown in Fig 3D on Janus particles. When the hydrophobic domains are small we hypothesize that surface-bound ‘half’ cavities (SBHC) form over them, in analogy to the bulk cavities discussed above. This is illustrated in the cartoon shown in Fig. 3C. Like cavities in bulk fluid, SBHCs result from a competition between bulk energies and interfacial penalties<sup>39</sup>. SBHCs lower the overall free energy of the entire system consisting of the NP, the interfacial water (i.e. water molecules close to the NP surface), even though they strain in the H-bond network surrounding the cavity, causing a corresponding increase in the *WSL*. In terms of number density of interfacial water molecules, cavitation will lead to an increase in the number density of water molecules over OT domains and a consequent necessary decrease in number density of water molecules over MHol domains. The results of molecular dynamics simulations on striped particles illustrated in Fig. 3 indeed confirm this trend: in Fig. 3H it is possible to observe how the number density of water molecules on OT domains increases as their width decreases with an opposite trend for the MH stripes (Fig. 3G, see also discussion below). We note that earlier simulations<sup>42</sup> of water confined between structured flat surfaces reported that hydrophobicity of a small hydrophobic patch, measured in terms of number density of water molecules above it, decreases when it is surrounded by hydrophilic molecules.

While cavitation is related mainly to the width of the hydrophobic stripes, the second mechanism we propose relates to the width of the hydrophilic stripes. We hypothesize that confinement of solvent molecules within a hydrophilic ‘groove’ of molecular dimensions, as

illustrated in Fig. 3C and E, is not favored entropically. This will be true for any solvent molecule inside a stripe ‘groove’ especially when the domain walls repel such molecules, and will be critically dependent on the size and the depth of the stripe. In Fig. 3A and C we illustrate this concept comparing the entropic freedom (represented by arrows) of two solvent molecules at identical distance from a wetting domain in the case of a flat local surface and that of a stripe groove. For narrow hydrophilic stripes, the effect of this ‘confinement mechanism’ is strong and becomes weaker as the hydrophilic stripes become wider. Wider hydrophilic stripes allow for more entropic freedom to the water molecules immediately over MHol and are expected to decrease the interfacial free energy and therefore decrease  $W_{SL}$ . A weak confinement means more entropic freedom for solvent molecules in the grooves, hence the water density in the grooves will also increase as MHol stripes become wider. Since these two molecular mechanisms – cavitation and confinement – have opposite influences on the  $W_{SL}$  with increasing MHol, they may lead to a net effect of non-monotonic dependence of  $W_{SL}$  on composition.

Our simulations confirm the presence of these two competing mechanisms. The inset of Fig. 3C and Fig. S9 show an example of H-bond bridges formed in cavitation. Figs. 3G and H show the number density of interfacial ( $< 1.3$  nm from NP surface) water molecules over MHol and OT domains,  $n_{MHol}$  and  $n_{OT}$ , respectively, as a function of composition in %MHol.  $n_{OT}$  (Fig. 3H) increases remarkably from 33% to 50% MHol and only slightly from 50% to 67% MHol, supporting our hypothesis that cavitation depends on OT stripe width and increases the water number density over OT end groups. As expected, from 33% to 50% MHol when cavitation is important, there is a decrease in  $n_{MHol}$  due to increasing cavitation. The confinement mechanism is confirmed by comparing  $n_{MHol}$  in the case of two particles (50% and 67% MHol) with almost the same hydrophobic stripe width and hence the same degree of cavitation.  $n_{MHol}$  changes between these two particles as confinement becomes more pronounced in the narrower stripes. A simplistic way of visualizing this mechanism is that every hydrophobic ‘wall’ (yellow cylinders in Fig. 3C) surrounding the hydrophilic grooves creates an excluded volume for the solvent (white regions in Fig. 3C); the narrower

the hydrophilic stripe (red cylinders in Fig. 3C), the larger the fraction of this excluded volume in the total groove volume. In this 50% to 67% MHol range where confinement is important, we see an increase in  $n_{\text{MHol}}$  (Fig. 3G) associated with increased entropic freedom of water molecules over MHol. Cavitation and confinement are therefore both present in the simulations and since they have opposing effects on  $WSL$  we expect that, in combination, they will lead to a non-monotonic dependence of  $WSL$  on %MHol.

Based on these findings, we explain the non-monotonic behavior of the nanoparticles'  $WSL$  as arising from a combination of cavitation and confinement of solvent molecules in the immediate neighborhood of the NP. We find that these two mechanisms both depend on stripe width, but oppose each other as the stripe width of the hydrophilic domain increases. The observed non-monotonic behavior is due to these two contrasting mechanisms (that act on the  $WSL$  per molecule) combined with the complex way in which domain size in these types of mixed SAMs evolves with composition. The result is a non-monotonic  $W_{SL}$  or  $\gamma_{SL}$  dependence on MHol content.

To further validate this interpretation of the data, we measured the effect of ionic strength on  $W_{SL}$ . It is known that salts raise the energy needed to form a cavity of a given size<sup>43</sup>. As a consequence the effect of cavitation energy in determining the  $W_{SL}$  should become less pronounced as salt concentration is increased, allowing its competing mechanism (confinement) to become more dominant. This implies that the deviations from the linear trend of  $W_{SL}$  would have one of two behaviors: either they decrease when large at zero salt concentration, or they increase when small at zero salt concentration. In the latter case, our interpretation implies that the contrasting effects balance out in de-ionized water, and that this balance is broken at high salt concentration. (At present we can only assume that our second effect, the solvent confinement effect, does not vary with ionic strength). We experimentally measured the  $W_{SL}$  for our surfaces in DI water and in water containing 3M NaCl. We plotted the deviations of  $W_{SL}$  from the expected linear trend in all cases. Figure 4 shows that the results support our hypothesis: when the deviations for DI water were large they become

small at the 3M concentration, and vice versa, corroborating the arguments developed through our simulations. It is worth discussing the case of the 50% composition. As discussed above, of the three types of particles studies with a dominant narrow stripe morphology, the 33% and the 67% compositions both showed strong deviations in DI water (20% and 9% respectively) but the 50% did not. However, measurements taken at 3M NaCl (where cavitation is attenuated) demonstrate a strong deviation (14%), pointing to the fact that both molecular mechanisms were strongly present at this composition but only coincidentally cancelling each other. This is to be expected once one accepts the two proposed mechanisms as the 50% composition is the one where both stripes are narrower and hence both effects should be strong.

In summary, our combined experimental and computational results show how the nanoscale structure of a surface affects the local organization of a wetting liquid in such a way as to appreciably modulate the interfacial solvent structure and energy. We see that this unanticipated and large effect is due both to the formation of cavities at the liquid-solid interface, and to confinement of solvent molecules over attractive domains. When combined, these two molecular-level mechanisms produce a non-monotonic wetting dependence on surface composition. This finding has important implications for biological processes where proteins and other biological structures exhibit nanoscale patterning and for our general understanding of solid-liquid interfaces. For instance, these findings show that concepts like hydrophobicity are ill-defined at the nanoscale (and hence possibly for most biological surfaces) where  $W_{SL}$  and thus contact angle measurements depend on nanostructure as much as on chemical composition.

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## Figure Legends:

**Figure 1:** Schematic drawing illustrating the fabrication of the different SAMs studied. Mixed ligand molecules (OT in yellow and MHol in red) are assembled either on flat gold surfaces (flat SAMs) or on gold nanoparticles (see SI for details). On flat SAMs, the ligand mixture separates in large ( $> 5$  nm across) homoligand domains. On nanoparticles, the ligands separate in stripe-like domains  $< 2$  nm wide. Nanoparticle films are then grown by stacking five layers of cross-linked nanoparticles (see SI for details).

**Figure 2:** SAM-AFM imaging of mixed and homogeneous SAMs in ultrapure water (A-C) and measurement of the corresponding  $W_{SL}$  by AFM and CA (D-E). For each image (A-C) the topographic image (left) and the corresponding phase image (right) are presented. The phase image is an indication of the energy dissipated by the vibrating tip in the surface solvation layers and consequently of  $W_{SL}$  (see SI). Darker regions in the phase images represent lower local energy dissipation by the tip and hence smaller  $W_{SL}$ . For flat SAMs (A), the phase image can distinguish between regions coated with the different ligands (OT is darker and MHol lighter, borderline highlighted). On the nanoparticles, mixed SAMs spontaneously phase-separate into local ribbon-like domains (arrows in B), also visible in the corresponding phase image where maxima in the local energy dissipation are visible (lighter regions). Single ligand molecules are resolved, showing structuring and contrast between the different ligand species (arrows). When imaging nanoparticles coated with one ligand, no clear ordering is visible on the nanoparticles (C), as confirmed in the phase image. The  $W_{SL}$  for flat SAMs (D) and coated nanoparticles (E) were obtained both from CA (blue) and AFM measurements (red) (see SI for details). Note the different scales for AFM and CA measurements. The variation of  $W_{SL}$  with ligand composition is clearly different for the flat SAMs and the nanoparticles, as highlighted by the grey shade in the background of the plots. Although AFM measurements present a large uncertainty for  $W_{SL}$ , there is good qualitative if not quantitative agreement with CA

measurements. In the case of flat SAMs (D), the AFM uncertainty is statistically increased by the relatively small sampling area (scan size) compared to the typical homoligand domain sizes. In (B) and (C), the left part of the picture is a simplified cartoon of the coated nanoparticle, serving as a guide for the eye. The contours of the cartoon nanoparticles have also been highlighted. The color scale (z-scale) is 1 nm, 4 nm and 3 nm for (A), (B) and (C) respectively and  $14^\circ$ ,  $40^\circ$  and  $40^\circ$  for the respective corresponding phase images.

**Figure 3:** Simulation results. (A) Schematic of water density profile and (D) 2D number density profile of water for a 50% MHol Janus nanoparticle with domain sizes commensurate to those of a flat SAM compared with (C) a schematic and (E) the water density profile for a striped 50% MHol nanoparticle. The fuzzy white regions in C are cavities bounded by H-bonded water molecules bridging over the OT domains. The fuzzy white region in A is an extended empty space over a large hydrophobic region. The hydrogen bond network of waters surrounding this region is irregular and broken. The inset in C shows h-bond bridges (see also Fig. SI9) anchored to MHol domains (red) and forming over OT domains (yellow). ((B) The region over an OT stripe from the Janus particle and the striped particle is highlighted to evince the role of cavitation in drawing water closer to the hydrophobic surface fraction. (F) Water number density profiles of striped nanoparticles with 33%, 50% and 67% MHol (left to right) composition as a function of radial distance,  $r$ , and angular coordinate,  $\theta$ . The alternating OT and MHol stripes are reflected as the peaks and troughs, respectively, in the density profile as  $\theta$  is varied for small  $r$ . (G,H) Number density of interfacial water molecules in nm<sup>-3</sup> over (G) MHol, nMHol, and (H) OT, nOT, domains, as a function of composition. The error bars reflect the standard deviation.

**Figure 4:** Plot of the deviation of  $W_{SL}$  as a function of composition from a linear trend. The deviation was calculated as the difference from the measured  $W_{SL}$  and the value expected from a linear trend found using the weighted average of the homoligand nanoparticle films. Increasing the salt concentration increases the work



to produce a cavity and decreases the likelihood of cavity formation. With the cavitation effect thus diminished, the balance with the confinement effect is shifted. The values found for the salt solution do not scale proportionally to the values obtained with pure water, indicating effects in addition to the modification of the surface tension of the water are occurring. The error bars reflect the standard deviation of the measurements per solution.

FIG 1

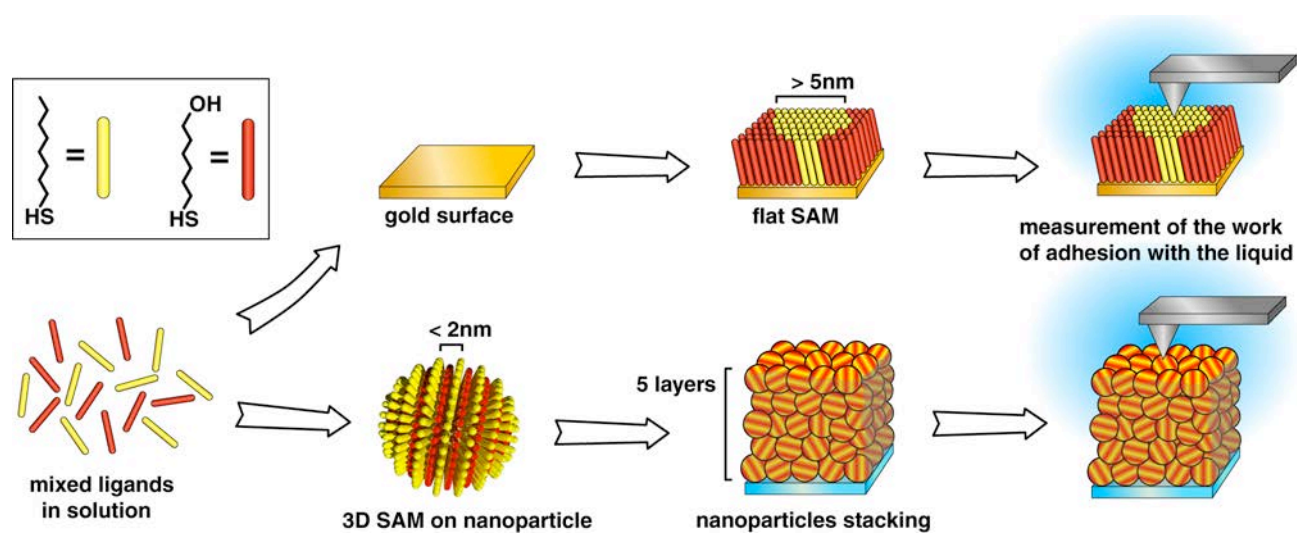


FIG 2

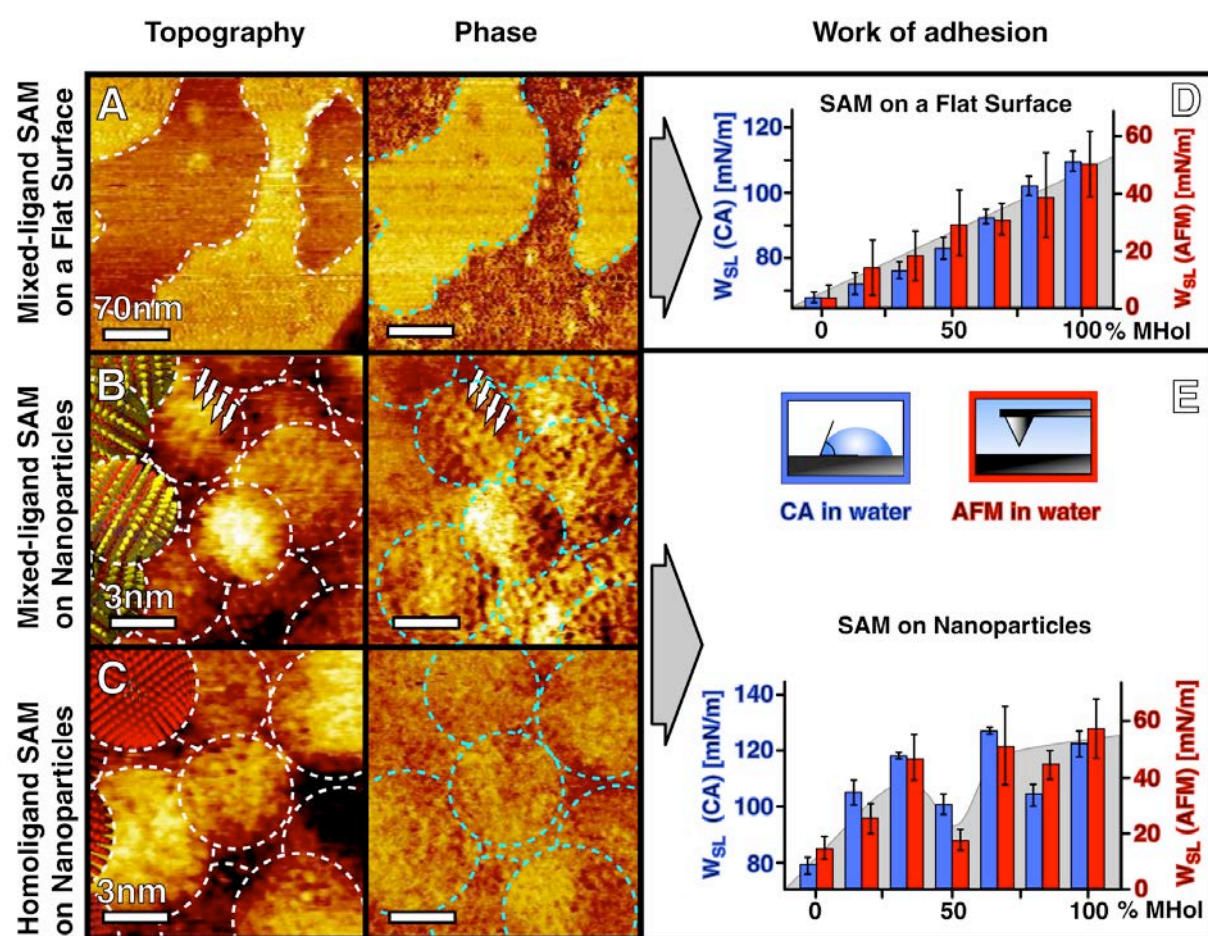


FIG 3

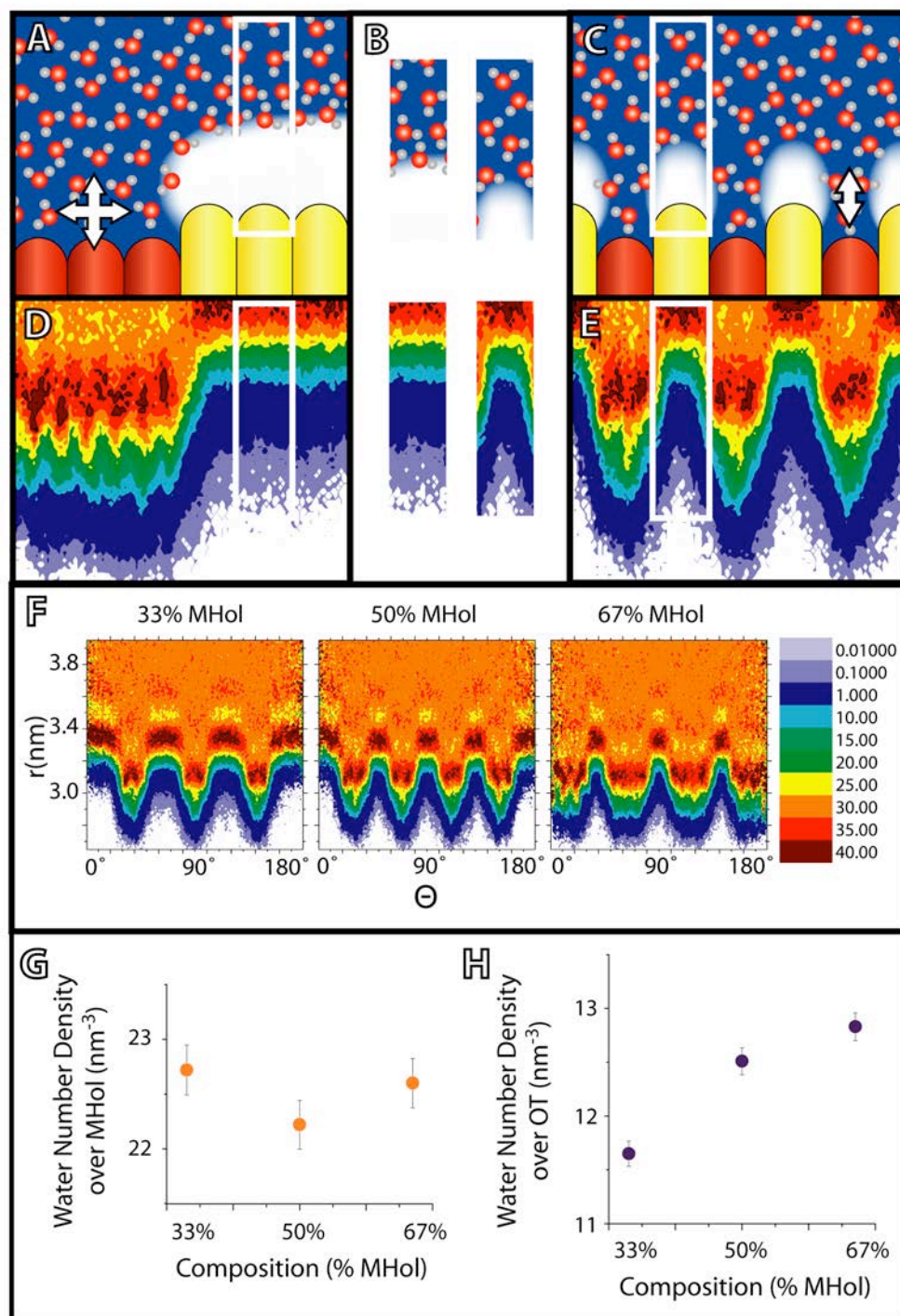


FIG 4

